

## Chemical Names and Formulas

### ❖ Key Vocabulary

**Monoatomic Ion** – Ions formed from a single atom

**Binary compound** – Compounds composed of two elements

**Nomenclature** – the “naming system” of binary ionic compounds

**Oxyanion** – polyatomic ions that contain oxygen

**Salt** – an ionic compound composed of a cation and the anion from an acid

### Chemical Formulas

Subscripts after an elemental symbol represent the number of atoms of the element in the compound

Molecular formulas represent how many atoms of each element are in a compound.  
( $C_2H_6$ ,  $H_2O$ ,  $C_6H_{12}O_6$ )

Ionic formulas show the simplest ratio, cations to anions, of atoms in the lattice

Parenthesis are used around a polyatomic ion in a chemical formula

Examples  $Al_2(SO_4)_3$ ,  $Cu(PO_4)_2$ ,  $Sr(NO_3)_2$

### Monoatomic Ions

– Ions formed from a single atom

Group 1 elements may lose one electron = 1+

Group 2 elements may lose two electrons = 2+

Group 13 elements may lose three electrons = 3+

Group 17 elements may gain one electron = 1-

Group 16 elements may gain two electrons = 2-

Group 15 elements may gain three electrons = 3-

Transition metals have various possibilities (no pattern)

### Naming monatomic Ions

Monatomic cations are simply called by the element's name

\*the stock system is used when there is more than one possibility for an element (more on this later)

Monatomic anions end in IDE such as Fluorine – Fluoride, Oxygen – Oxide  
(reference table 1 on page 221)

### Binary Ionic Compounds

1. Write the symbols for the ions with the cation first ( $Al^{3+}$   $O^{2-}$ )

2. Cross over the charges writing the charge of one ion as the subscript for the other ion ( $Al_2^{3+}$   $O_3^{2-}$ )

3. Multiply subscript and superscript for each ion then find the lowest whole number ratio ( $Al_2O_3$ )

**Name it!!! Aluminum Oxide**

Examples:  $Zn^{2+}$   $I^-$        $Zn^{2+}$   $S^{2-}$

(More on p 223 practice problems)

# Chemical Names And Formulas Study Guide

**James E. Brady, Fred Senese**



## Chemical Names And Formulas Study Guide:

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Katherine J. Denniston, Denniston Katherine, 2006     *Journal of Chemical Education* ,1925 Includes Report of New England Association of Chemistry Teachers and Proceedings of the Pacific Southwest Association of Chemistry Teachers

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