

CHEM 110
Extra In-Class Practice Problems
Lecture Set 3

1. Answer the following questions dealing with Avogadro's number.

A. How many molecules present in 2.0 moles CH_4 ?

$$\frac{2.0 \text{ mol } \text{CH}_4}{1 \text{ mol } \text{CH}_4} \times \frac{6.022 \times 10^{23} \text{ molecules}}{1 \text{ mol } \text{CH}_4} = 1.2 \times 10^{24} \text{ molecules } \text{CH}_4$$

How many atoms are present?

$$\frac{1.2 \times 10^{24} \text{ molecules } \text{CH}_4}{1 \text{ molecule } \text{CH}_4} \times \frac{5 \text{ atoms}}{1 \text{ molecule } \text{CH}_4} = 6.0 \times 10^{24} \text{ atoms}$$

B. How many formula units of CaCl_2 are present in 0.543 mol CaCl_2 ?

$$\frac{0.543 \text{ mol } \text{CaCl}_2}{1 \text{ mol } \text{CaCl}_2} \times \frac{6.022 \times 10^{23} \text{ f.u.}}{1 \text{ mol } \text{CaCl}_2} = 3.27 \times 10^{23} \text{ f.u. } \text{CaCl}_2$$

C. 9.81×10^{21} molecules O_2 is how many moles?

$$\frac{9.81 \times 10^{21} \text{ molecules } \text{O}_2}{6.022 \times 10^{23} \text{ molecules}} \times \frac{1 \text{ mol } \text{O}_2}{6.022 \times 10^{23} \text{ molecules}} = 0.0163 \text{ mol } \text{O}_2$$

D. How many atoms are present in 0.371 mol P_2O_5 ?

$$\frac{0.371 \text{ mol } \text{P}_2\text{O}_5}{1 \text{ mol } \text{P}_2\text{O}_5} \times \frac{6.022 \times 10^{23} \text{ molecules } \text{P}_2\text{O}_5}{1 \text{ mol } \text{P}_2\text{O}_5} \times \frac{7 \text{ atoms}}{1 \text{ molecule } \text{P}_2\text{O}_5} = 1.56 \times 10^{24} \text{ atoms}$$

Chemistry 110 Practice Problems

**David Mautner Himmelblau, James B.
Riggs**



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Safety),2010-09-14 This Guidelines book provides technical information on how to conduct a consequence analysis to satisfy your company's needs and the EPA rules It covers quantifying the size of a release dispersion of vapor clouds to an endpoint concentration outcomes for various types of explosions and fires and the effect of the release on people and structures Special Details Includes CD ROM with example problems worked using Excel and Quattro Pro For use with Windows 95 98 and NT

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Basic Principles and Calculations in Chemical Engineering David Mautner Himmelblau,James B. Riggs,2012 Best selling introductory chemical engineering book now updated with far more coverage of biotech nanotech and green engineering Thoroughly covers material balances gases liquids and energy balances Contains new biotech and bioengineering problems throughout

Partial Solutions Manual Darrell D. Ebbing,David Bookin,1998 Register of the University of California University of California, Berkeley,1928 General Catalogue Massachusetts Institute of Technology,1925 **Annual Catalogue** Massachusetts Institute of Technology,1916

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