

CHM 104 Chapter 16 Pretest Questions (Acids and Bases) with key

Indicate the answer choice that best completes the statement or answers the question.

1. What is the H_3O^+ concentration of 0.0044 M $\text{KOH}(\text{aq})$ at 25 °C? ($K_w = 1.01 \times 10^{-14}$)
 - a. 2.3×10^{-12} M
 - b. 4.4×10^{-3} M
 - c. 1.0×10^{-14} M
 - d. 1.0×10^{-7} M
 - e. 4.4×10^{-17} M

3. What is the conjugate base of H_3PO_4 in aqueous solution?
 - a. not possible
 - b. PO_4^{3-}
 - c. HPO_4^{2-}
 - d. H_2PO_4^-
 - e. H_3PO_4

5. What is the pH of the solution which results from mixing 75 mL of 0.50 M $\text{NH}_3(\text{aq})$ and 75 mL of 0.50 M $\text{HCl}(\text{aq})$ at 25 °C? (K_b for $\text{NH}_3 = 1.8 \times 10^{-5}$)
 - a. 0.60
 - b. 2.67
 - c. 4.74
 - d. 4.93
 - e. 9.26

7. Which of the following species is the strongest acid in an aqueous solution?
 - a. $\text{CH}_3\text{CH}_2\text{CO}_2\text{H}$
 - b. $\text{CH}_2\text{ClCO}_2\text{H}$
 - c. $\text{CH}_3\text{CO}_2\text{H}$
 - d. $\text{CCl}_3\text{CO}_2\text{H}$
 - e. $\text{CHCl}_2\text{CO}_2\text{H}$

9. What is the conjugate base of $[\text{Fe}(\text{H}_2\text{O})_6]^{3+}(\text{aq})$?
 - a. H_3O^+
 - b. $[\text{Fe}(\text{H}_2\text{O})_6]^{2+}$
 - c. $[\text{Fe}(\text{H}_2\text{O})_5\text{H}_2\text{O}]^{4+}$
 - d. $[\text{Fe}(\text{H}_2\text{O})_5\text{OH}]^{2+}$
 - e. $[\text{Fe}(\text{H}_2\text{O})_5]^{3+}$

11. Which equation depicts hydrogen phosphate ion behaving as a Bronsted-Lowry base in water?
 - a. $\text{HPO}_4^{2-}(\text{aq}) + \text{H}_2\text{O}(\ell) \rightleftharpoons \text{H}_2\text{PO}_4^-(\text{aq}) + \text{OH}^-(\text{aq})$
 - b. $\text{HPO}_4^{2-}(\text{aq}) + \text{OH}^-(\text{aq}) \rightleftharpoons \text{PO}_4^{3-}(\text{aq}) + \text{H}_2\text{O}(\ell)$
 - c. $\text{HPO}_4^{2-}(\text{aq}) + \text{H}_2\text{O}(\ell) \rightleftharpoons \text{PO}_4^{3-}(\text{aq}) + \text{H}_3\text{O}^+(\text{aq})$

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