

## CHM 104 Chapter 16 Pretest Questions (Acids and Bases) with key

Indicate the answer choice that best completes the statement or answers the question.

1. What is the  $\text{H}_3\text{O}^+$  concentration of 0.0044 M  $\text{KOH}(\text{aq})$  at 25  $^{\circ}\text{C}$ ? ( $K_w = 1.01 \times 10^{-14}$ )

- a.  $2.3 \times 10^{-12} \text{ M}$
- b.  $4.4 \times 10^{-3} \text{ M}$
- c.  $1.0 \times 10^{-14} \text{ M}$
- d.  $1.0 \times 10^{-7} \text{ M}$
- e.  $4.4 \times 10^{-17} \text{ M}$

3. What is the conjugate base of  $\text{H}_3\text{PO}_4$  in aqueous solution?

- a. not possible
- b.  $\text{PO}_4^{3-}$
- c.  $\text{HPO}_4^{2-}$
- d.  $\text{H}_2\text{PO}_4^-$
- e.  $\text{H}_3\text{PO}_4$

5. What is the pH of the solution which results from mixing 75 mL of 0.50 M  $\text{NH}_3(\text{aq})$  and 75 mL of 0.50 M  $\text{HCl}(\text{aq})$  at 25  $^{\circ}\text{C}$ ? ( $K_b$  for  $\text{NH}_3 = 1.8 \times 10^{-5}$ )

- a. 0.60
- b. 2.67
- c. 4.74
- d. 4.93
- e. 9.26

7. Which of the following species is the strongest acid in an aqueous solution?

- a.  $\text{CH}_3\text{CH}_2\text{CO}_2\text{H}$
- b.  $\text{CH}_2\text{ClCO}_2\text{H}$
- c.  $\text{CH}_3\text{CO}_2\text{H}$
- d.  $\text{CCl}_3\text{CO}_2\text{H}$
- e.  $\text{CHCl}_2\text{CO}_2\text{H}$

9. What is the conjugate base of  $[\text{Fe}(\text{H}_2\text{O})_6]^{3+}(\text{aq})$ ?

- a.  $\text{H}_3\text{O}^+$
- b.  $[\text{Fe}(\text{H}_2\text{O})_6]^{2+}$
- c.  $[\text{Fe}(\text{H}_2\text{O})_5\text{H}_2\text{O}]^{4+}$
- d.  $[\text{Fe}(\text{H}_2\text{O})_5\text{OH}]^{2+}$
- e.  $[\text{Fe}(\text{H}_2\text{O})_5]^{3+}$

11. Which equation depicts hydrogen phosphate ion behaving as a Brønsted-Lowry base in water?

- a.  $\text{HPO}_4^{2-}(\text{aq}) + \text{H}_2\text{O}(\ell) \rightleftharpoons \text{H}_2\text{PO}_4^-(\text{aq}) + \text{OH}^-(\text{aq})$
- b.  $\text{HPO}_4^{2-}(\text{aq}) + \text{OH}^-(\text{aq}) \rightleftharpoons \text{PO}_4^{3-}(\text{aq}) + \text{H}_2\text{O}(\ell)$
- c.  $\text{HPO}_4^{2-}(\text{aq}) + \text{H}_2\text{O}(\ell) \rightleftharpoons \text{PO}_4^{3-}(\text{aq}) + \text{H}_3\text{O}^+(\text{aq})$

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