

EQUILIBRIUM CHEAT SHEET

Character of equilibrium: equilibrium, equilibrium, whether the rate of the forward reaction equals the rate of the reverse reaction. Equilibrium is a dynamic state of temporary balance of forward and reverse reactions.

Nonsolvent Equilibrium	
<p>Reactions in a nonreactive medium, different solvents</p> <p>When concentrations of reactants are the same, the forward reaction rate equals the reverse reaction rate, and the equilibrium constant of reaction is 1.0 (no net reaction or change).</p>	
<p>At 1 mole/liter, equilibrium of solvents with different solvents: when the forward reaction rate equals the reverse reaction rate, then the equilibrium constant of reaction is 1.0 (no net reaction or change).</p>	
<p>Equilibrium Constant Formula: $K_{eq} = \frac{[A][B]}{[C][D]} = 1.0 = 1.0$</p>	

Equilibrium of solvents, K_{sol}	
<p>When solvents are different, the equilibrium constant of reaction is not 1.0 (no net reaction or change). The equilibrium constant is a ratio of the concentrations of products to reactants when the reaction is at equilibrium.</p>	
$K_{sol} = \frac{[A][B]}{[C][D]} \neq 1.0 \neq 1.0$	
<p>Equilibrium Constant Expression:</p> $K_{sol} = \frac{[A]^x[B]^y}{[C]^p[D]^q}$ <p>The greater the equilibrium constant, the greater the reaction rate and the greater the extent of reaction.</p>	

Reaction Quotient, Q	
<p>Reaction Quotient, Q, is a ratio of the concentrations of reactants to the concentrations of products at any time during the reaction. The reaction quotient equals 1.0 at the point of equilibrium or equilibrium is reached when the reaction quotient is 1.0, and reaction is at equilibrium. The reaction is at equilibrium when the reaction quotient is 1.0.</p>	
$Q = \frac{[A]^x[B]^y}{[C]^p[D]^q}$	

Equilibrium Position	
<p>When $K_{eq} > 1$</p>	
<p>Forward reaction is favored, equilibrium is shifted to products.</p>	
<p>When $K_{eq} < 1$</p>	
<p>Reverse reaction is favored, equilibrium is shifted to reactants.</p>	
<p>When $K_{eq} = 1$</p>	
<p>Forward and reverse reaction rates are equal, equilibrium is at a standstill.</p>	



K-equilibrium & chemical Equilibrium	
Equilibrium Constant	Definition
$K_{eq} = 1$	Slow conversion of reactants, equilibrium is not shifted to products.
$K_{eq} < 1$	Slow reaction, slow forward reaction is favored.
$K_{eq} > 1$	Fast reaction, fast forward reaction is favored.

The Equilibrium Constant vs. Equilibrium Constant	
Equilibrium Constant vs. Equilibrium Constant	Definition
$Q = K_{eq}$	Reaction is at equilibrium: $2\text{NO}_2 \rightleftharpoons 2\text{NO} + \text{O}_2$
$Q < K_{eq}$	Reaction is slow, forward reaction is favored, reaction is not at equilibrium: $2\text{NO}_2 \rightleftharpoons 2\text{NO} + \text{O}_2$
$Q > K_{eq}$	Reaction is slow, reverse reaction is favored, reaction is not at equilibrium: $2\text{NO}_2 \rightleftharpoons 2\text{NO} + \text{O}_2$

The Classification Principle: The nature of the reaction, equilibrium and equilibrium constant.		
Classification	Nature of Change	Description
Homogeneous	<ul style="list-style-type: none"> Reactants and products are in the same phase Reactants and products are in different phases 	$\text{NO}_2 + \text{NO}_2 \rightleftharpoons 2\text{NO}_2$ Reactants and products are in the same phase Reactants and products are in different phases
Heterogeneous	<ul style="list-style-type: none"> Reactants and products are in different phases Reactants and products are in the same phase Reactants and products are in different phases 	$\text{2NO}_2 \rightleftharpoons \text{2NO} + \text{O}_2$ Reactants and products are in different phases Reactants and products are in the same phase Reactants and products are in different phases
Heterogeneous	<ul style="list-style-type: none"> Reactants and products are in the same phase Reactants and products are in different phases 	$\text{2NO}_2 \rightleftharpoons \text{2NO} + \text{O}_2$ Reactants and products are in the same phase Reactants and products are in different phases

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