

EQUILIBRIUM CHEAT SHEET

Chemical equilibrium, dynamic equilibrium, when the rate of the forward reaction equals the rate of the reverse reaction. Equilibrium reactions are represented by this symbol: \rightleftharpoons

What is Equilibrium?

Equilibrium is a constant concentration of reactants and products

When a reversible reaction takes place in a closed system, the forward and reverse reactions occur at the same rate. The system is called dynamic equilibrium if the concentrations of reactants and products do not change over time.

At equilibrium, the concentrations of reactants and products are constant. The system is called dynamic equilibrium because the forward and reverse reactions are still occurring, but the overall concentrations remain the same.

Equilibrium Constant Formulas

$$K_c = \frac{[C]^c [D]^d}{[A]^a [B]^b}$$

Equilibrium Constant, K_c

The equilibrium constant, K_c , is a ratio of the concentrations of products to reactants, raised to the power of their stoichiometric coefficients, at equilibrium. It is a constant for a given reaction at a given temperature.

$$K_c = \frac{[C]^c [D]^d}{[A]^a [B]^b}$$

Equilibrium Constant Equation

$$K_c = \frac{[C]^c [D]^d}{[A]^a [B]^b}$$

The equilibrium constant, K_c , is a ratio of the concentrations of products to reactants, raised to the power of their stoichiometric coefficients, at equilibrium. It is a constant for a given reaction at a given temperature.

Reaction Quotient, Q

The reaction quotient, Q , is a ratio of the concentrations of products to reactants, raised to the power of their stoichiometric coefficients, at any point in time. It is used to determine if a system is at equilibrium. If $Q = K_c$, the system is at equilibrium. If $Q < K_c$, the system is not at equilibrium and the forward reaction is favored. If $Q > K_c$, the system is not at equilibrium and the reverse reaction is favored.

$$Q = \frac{[C]^c [D]^d}{[A]^a [B]^b}$$

Equilibrium Constants

$$K_c = \frac{[C]^c [D]^d}{[A]^a [B]^b}$$

Example

$$2 H_2 + O_2 \rightleftharpoons 2 H_2O$$

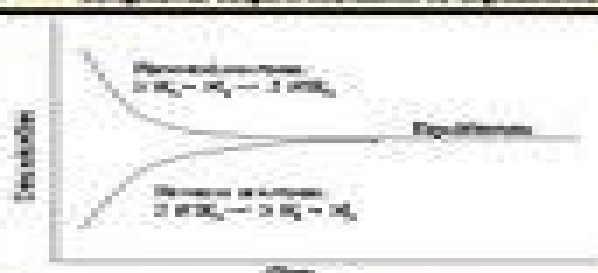
Forward Reaction

$$2 H_2 + O_2 \rightarrow 2 H_2O$$

Reverse Reaction

$$2 H_2O \rightarrow 2 H_2 + O_2$$

Graphical Representation of Equilibrium



Equilibrium Constant Relationship

Equilibrium Constant Value	Meaning
$K_c > 1$	Concentration of products is greater than concentration of reactants
$K_c < 1$	Concentration of reactants is greater than concentration of products
$K_c = 1$	Concentration of reactants is equal to concentration of products

Reaction Quotient vs. Equilibrium Constant

Equilibrium Constant vs. Reaction Quotient	Meaning
$Q = K_c$	System is at equilibrium. Forward reaction: $2 H_2 + O_2 \rightleftharpoons 2 H_2O$
$Q < K_c$	Forward reaction is favored. Forward reaction: $2 H_2 + O_2 \rightarrow 2 H_2O$
$Q > K_c$	Reverse reaction is favored. Reverse reaction: $2 H_2O \rightarrow 2 H_2 + O_2$

Le Chatelier's Principle: The effect of changing conditions on equilibrium

Condition	Effect of Change	Example
Concentration	<ul style="list-style-type: none"> Increasing concentration of reactants shifts equilibrium to the right. Increasing concentration of products shifts equilibrium to the left. 	$2 H_2 + O_2 \rightleftharpoons 2 H_2O$ Increasing concentration of H_2 or O_2 shifts equilibrium to the right, increasing concentration of H_2O .
Pressure	Increasing pressure shifts equilibrium to the side with fewer moles of gas.	$2 H_2 + O_2 \rightleftharpoons 2 H_2O$ Increasing pressure shifts equilibrium to the right, increasing concentration of H_2O .
Temperature	Increasing temperature shifts equilibrium to the side with higher energy (endothermic reaction).	$2 H_2 + O_2 \rightleftharpoons 2 H_2O$ Increasing temperature shifts equilibrium to the left, decreasing concentration of H_2O .

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Table of Contents Chemistry Chemical Equilibrium Study Guide Answers

1. Understanding the eBook Chemistry Chemical Equilibrium Study Guide Answers
 - The Rise of Digital Reading Chemistry Chemical Equilibrium Study Guide Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Chemical Equilibrium Study Guide Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chemistry Chemical Equilibrium Study Guide Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry Chemical Equilibrium Study Guide Answers
 - Personalized Recommendations
 - Chemistry Chemical Equilibrium Study Guide Answers User Reviews and Ratings

- Chemistry Chemical Equilibrium Study Guide Answers and Bestseller Lists
- 5. Accessing Chemistry Chemical Equilibrium Study Guide Answers Free and Paid eBooks
 - Chemistry Chemical Equilibrium Study Guide Answers Public Domain eBooks
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 - Adjustable Fonts and Text Sizes of Chemistry Chemical Equilibrium Study Guide Answers
 - Highlighting and Note-Taking Chemistry Chemical Equilibrium Study Guide Answers
 - Interactive Elements Chemistry Chemical Equilibrium Study Guide Answers
- 8. Staying Engaged with Chemistry Chemical Equilibrium Study Guide Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chemistry Chemical Equilibrium Study Guide Answers
- 9. Balancing eBooks and Physical Books Chemistry Chemical Equilibrium Study Guide Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemistry Chemical Equilibrium Study Guide Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Chemistry Chemical Equilibrium Study Guide Answers
 - Setting Reading Goals Chemistry Chemical Equilibrium Study Guide Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chemistry Chemical Equilibrium Study Guide Answers
 - Fact-Checking eBook Content of Chemistry Chemical Equilibrium Study Guide Answers
 - Distinguishing Credible Sources

13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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