

Chemistry

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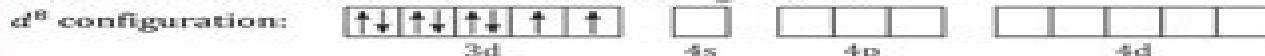
(Chapter 5) (Coordination Compound) (Intext Questions)
(Class 12)

Question 5.5:

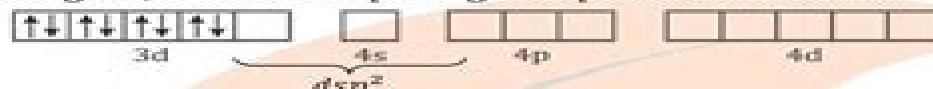
Explain on the basis of valence bond theory that $[\text{Ni}(\text{CN})_4]^{2-}$ ion with square planar structure is diamagnetic and the $[\text{NiCl}_4]^{2-}$ ion with tetrahedral geometry is paramagnetic.

Answer 5.5:

Ni is in the +2 oxidation state i.e., in d^8 configuration.



There are 4 CN^- ions. Thus, it can either have a tetrahedral geometry or square planar geometry. Since CN^- ion is a strong field ligand, it causes the pairing of unpaired 3d electrons.



It now undergoes dsp^2 hybridization. Since all electrons are paired, it is diamagnetic. In case of $[\text{NiCl}_4]^{2-}$, Cl^- ion is a weak field ligand. Therefore, it does not lead to the pairing of unpaired 3d electrons. Therefore, it undergoes sp^3 hybridization.



Since there are 2 unpaired electrons in this case, it is paramagnetic in nature.

Question 5.6:

$[\text{NiCl}_4]^{2-}$ is paramagnetic while $[\text{Ni}(\text{CO})_4]$ is diamagnetic though both are tetrahedral. Why?

Answer 5.6:

Though both $[\text{NiCl}_4]^{2-}$ and $[\text{Ni}(\text{CO})_4]$ are tetrahedral, their magnetic characters are different. This is due to a difference in the nature of ligands. Cl^- is a weak field ligand and it does not cause the pairing of unpaired 3d electrons. Hence, $[\text{NiCl}_4]^{2-}$ is paramagnetic.



In $[\text{Ni}(\text{CO})_4]$, Ni is in the zero oxidation state i.e., it has a configuration of $3d^8 4s^2$.



But CO is a strong field ligand. Therefore, it causes the pairing of unpaired 3d electrons. Also, it causes the 4s electrons to shift to the 3d orbital, thereby giving rise to sp^3 hybridization. Since no unpaired electrons are present in this case, $[\text{Ni}(\text{CO})_4]$ is diamagnetic.

Question 5.7:

$[\text{Fe}(\text{H}_2\text{O})_6]^{3+}$ is strongly paramagnetic whereas $[\text{Fe}(\text{CN})_6]^{3-}$ is weakly paramagnetic. Explain.

Answer 5.7:

In both $[\text{Fe}(\text{H}_2\text{O})_6]^{3+}$ and $[\text{Fe}(\text{CN})_6]^{3-}$, Fe exists in the +3 oxidation state i.e., in d^5 configuration.



Since CN^- is a strong field ligand, it causes the pairing of unpaired electrons. Therefore, there is only one unpaired electron left in the d-orbital.



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Table of Contents Chemistry Class Xii Guide Answers

1. Understanding the eBook Chemistry Class Xii Guide Answers
 - The Rise of Digital Reading Chemistry Class Xii Guide Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Class Xii Guide Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chemistry Class Xii Guide Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry Class Xii Guide Answers
 - Personalized Recommendations

- Chemistry Class Xii Guide Answers User Reviews and Ratings
- Chemistry Class Xii Guide Answers and Bestseller Lists

5. Accessing Chemistry Class Xii Guide Answers Free and Paid eBooks

- Chemistry Class Xii Guide Answers Public Domain eBooks
- Chemistry Class Xii Guide Answers eBook Subscription Services
- Chemistry Class Xii Guide Answers Budget-Friendly Options

6. Navigating Chemistry Class Xii Guide Answers eBook Formats

- ePUB, PDF, MOBI, and More
- Chemistry Class Xii Guide Answers Compatibility with Devices
- Chemistry Class Xii Guide Answers Enhanced eBook Features

7. Enhancing Your Reading Experience

- Adjustable Fonts and Text Sizes of Chemistry Class Xii Guide Answers
- Highlighting and Note-Taking Chemistry Class Xii Guide Answers
- Interactive Elements Chemistry Class Xii Guide Answers

8. Staying Engaged with Chemistry Class Xii Guide Answers

- Joining Online Reading Communities
- Participating in Virtual Book Clubs
- Following Authors and Publishers Chemistry Class Xii Guide Answers

9. Balancing eBooks and Physical Books Chemistry Class Xii Guide Answers

- Benefits of a Digital Library
- Creating a Diverse Reading Collection Chemistry Class Xii Guide Answers

10. Overcoming Reading Challenges

- Dealing with Digital Eye Strain
- Minimizing Distractions
- Managing Screen Time

11. Cultivating a Reading Routine Chemistry Class Xii Guide Answers

- Setting Reading Goals Chemistry Class Xii Guide Answers
- Carving Out Dedicated Reading Time

12. Sourcing Reliable Information of Chemistry Class Xii Guide Answers

- Fact-Checking eBook Content of Chemistry Class Xii Guide Answers

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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