

Chemistry Formulas

Dilution	$M_1V_1 = M_2V_2$	$pK_a \& pK_b$	$pK_a = -\log[K_a]$
Energy of a photon	$E_{\text{photon}} = hf = hc/\lambda$	$pK_a \& pK_b$	$pK_b = -\log[K_b]$
Pressure	$P = F/A$	$pK_a \& pK_b$	$pK_a + pK_b = 14$
Average Kinetic Energy	$KE_{\text{avg}} = \frac{3}{2}RT$	Neutralization Reaction	$n_A M_A V_A = n_B M_B V_B$
Ideal Gas Law	$PV = nRT$	Buffers	$pH = pK_a + \log [A^-] [HA]$
Combined Gas Law	$\frac{P_1V_1}{n_1T_1} = \frac{P_2V_2}{n_2T_2}$	Enthalpy of Formation	$\Delta H_f = \Sigma n\Delta H_f^{\circ}(\text{product}) - \Sigma n\Delta H_f^{\circ}(\text{reactants})$
Dalton's Law of Partial Pressures	$P_{\text{total}} = P_A + P_B + \dots$	First Law of Thermodynamics	$dE = q + w$
Dalton's Law of Partial Pressures	$P_A = X_A P_{\text{total}}$	Pressure Volume Work	$w = -P\Delta V$
Graham's Law of Effusion	$\frac{r_{\text{A}}}{r_{\text{B}}} = \sqrt{\frac{M_{\text{B}}}{M_{\text{A}}}}$	Calorimetry Thermal Energy (q)	$q = -C_{\text{calorimeter}} \Delta T$
Henry's Law	$P_A = k_H[A]$	Heat Curves & Thermal Energy (q)	$q = mc\Delta T$
Freezing Point Depression	$\Delta T_f = -iK_f M$	Heat Curves & Thermal Energy (q)	$q = m\Delta H_f^{\circ}$
Boiling Point Elevation	$\Delta T_b = iK_b m$	Heat Curves & Thermal Energy (q)	$q = m\Delta H_{\text{vaporization}}$
Vapor Pressure Depression (Raoult's Law)	$P_{\text{total}} = X_{\text{water}} P^{\circ}_{\text{water}}$	Entropy (S)	$\Delta S = \Sigma nS_{\text{products}} - \Sigma nS_{\text{reactants}}$
Osmotic Pressure (π)	$\pi = iMRT$	Bond Dissociation Energy	$\Delta H = \Sigma n\Delta H_{\text{reactants}} - \Sigma n\Delta H_{\text{products}}$
Arrhenius Equation	$k = Ae^{-\nu\epsilon/RT}$	Gibb's Free Energy (ΔG)	$\Delta G^{\circ} = \Delta H^{\circ} - T\Delta S^{\circ}$
Equilibrium Constant Expressions	$K_c = \frac{[\text{products}]}{[\text{reactants}]}$	Gibb's Free Energy (ΔG)	$\Delta G = \Delta G^{\circ} + RT\ln Q$
Equilibrium Constant Expressions	$K_{\text{eq}} = \frac{k_{\text{forward}}}{k_{\text{reverse}}}$	Gibb's Free Energy (ΔG)	$\Delta G^{\circ} = -RT^2 \ln(K_{\text{eq}})$
Equilibrium Constant Expressions	$K_P = \frac{P_{\text{products}}}{P_{\text{reactants}}}$	Standard Cell Potential	$E^{\circ} = E^{\circ}\text{reduction} + E^{\circ}\text{oxidation}$
Reaction Quotient (Q)	$Q = \frac{[\text{products}]}{[\text{reactants}]}$	Standard Cell Potential	$E^{\circ} = E^{\circ}\text{cathode} + E^{\circ}\text{anode}$
Solubility Product Constant (K_{sp})	$K_{sp} = \frac{[\text{products}]}{[\text{reactants}]}$	Nernst Equation	$E_{\text{cell}} = E^{\circ} - \frac{0.0592}{n} \log Q$
pH & pOH	$pH = -\log[H^+]$	Kinetics (always 1st order)	$N = N_0 e^{-kt}$
pH & pOH	$pOH = -\log[OH^-]$	Kinetics (always 1st order)	$\ln N = \ln N_0 - kt$
pH & pOH	$pH + pOH = 14$	Kinetics (always 1st order)	$t_{1/2} = \frac{0.693}{k}$
		Nuclear Binding Energy	$E = \Delta mc^2$

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Ettore Molinari

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