

1. How many atoms are in 4 moles of iron?

$$4 \times 6.02 \times 10^{23} = 2.4 \times 10^{24}$$

2. What is the percentage by mass of oxygen in water?

$$\frac{16}{18} \times 100 = 89\%$$



$$H_2 = 2 \times 1 = 2$$

$$O = 1 \times 16 = 16$$

3. A compound is 57.5 % sodium, 40% oxygen and 2.5 % hydrogen. What is its empirical formula?

$$57.5g \times \frac{1 \text{ mol}}{23g \text{ Na}} = 2.5 \text{ mol Na}$$

$$2.5g \text{ H} \times \frac{1 \text{ mol}}{1g \text{ H}} = 2.5 \text{ mol H}$$

$$40g \times \frac{1 \text{ mol}}{16g \text{ O}} = 2.5 \text{ mol O}$$

1:1:1 ratio



4. A compound with an empirical formula of  $CH_2$  has a molecular weight of 42. What is its molecular formula?



$$C = 3 \times 12 = 36$$

$$H = 6 \times 1 = 6$$

$$42$$

5. At STP ~~48~~ 44.8 liters of neon would contain the same number of molecules as how many grams of Argon? What volume would the argon occupy?

$$44.8 \text{ L Ne} \times \frac{1 \text{ mol Ne}}{22.4 \text{ L Ne}} = 2 \text{ mol Ne}$$

$$2 \text{ mol Ar} \times \frac{39.9g \text{ Ar}}{1 \text{ mol Ar}} = 79.8g \text{ Ar}$$

$$\text{volume of Ar} = 44.8 \text{ L}$$

6. Sodium and chlorine react to form sodium chloride. If 4 moles of chlorine react, how many moles of sodium chloride will form?



$$4 \text{ mol } Cl_2 \times \frac{2 \text{ mol Na}}{1 \text{ mol } Cl_2} = 8 \text{ mol Na}$$

7. How many liters of chlorine at STP would be needed to react with 92 grams of sodium?

$$92g \text{ Na} \times \frac{1 \text{ mol Na}}{23g \text{ Na}} \times \frac{1 \text{ mol } Cl_2}{2 \text{ mol Na}} \times \frac{22.4 \text{ L}}{1 \text{ mol } Cl_2} = 44.8 \text{ L } Cl_2$$

8. How many milliliters of chlorine were required in question 7?

$$44.8 \text{ L} \times \frac{1000 \text{ mL}}{1 \text{ L}} = 44,800 \text{ mL } Cl_2$$

9. Write and balance the equation for the combustion of butane.



10. If 232.0 grams of butane are combined with 192 grams of oxygen and a reaction occurs, how many grams of water will form?

$$232g \text{ C}_4\text{H}_{10} \times \frac{1 \text{ mol C}_4\text{H}_{10}}{58g \text{ C}_4\text{H}_{10}} \times \frac{10 \text{ mol H}_2\text{O}}{2 \text{ mol C}_4\text{H}_{10}} \times \frac{18g \text{ H}_2\text{O}}{1 \text{ mol H}_2\text{O}} = 360g \text{ H}_2\text{O}$$

$$192g \text{ O}_2 \times \frac{1 \text{ mol O}_2}{32g \text{ O}_2} \times \frac{10 \text{ mol H}_2\text{O}}{13 \text{ mol O}_2} \times \frac{18g \text{ H}_2\text{O}}{1 \text{ mol H}_2\text{O}} = \boxed{83g \text{ H}_2\text{O}}$$

11. If the reaction in number 9 was performed and 2 liters of carbon dioxide were produced at STP, what was the percent yield?

$$2 \text{ L } CO_2 \times \frac{1 \text{ mol } CO_2}{22.4 \text{ L } CO_2} \times \frac{10 \text{ mol H}_2\text{O}}{8 \text{ mol } CO_2} \times \frac{18g \text{ H}_2\text{O}}{1 \text{ mol H}_2\text{O}} = 2g \text{ H}_2\text{O}$$

$$\frac{2g \text{ H}_2\text{O}}{83g \text{ H}_2\text{O}} \times 100 = \boxed{2.4\%}$$

12. Write and balance the equation for the reaction of aluminum with hydrochloric acid.



see

$$\frac{23g \text{ H}_2}{2g} \times \frac{3 \text{ mol H}_2}{2 \text{ mol Al}} \times \frac{27g \text{ Al}}{1 \text{ mol}} = 18g$$

$$\frac{18g}{76} \times 100 = 24\%$$

13. If you are given 108 grams of aluminum to react with 146 grams of HCl, what will be the limiting reactant?

$$108g \text{ Al} \times \frac{1 \text{ mol Al}}{26.98g \text{ Al}} \times \frac{3 \text{ mol H}_2}{2 \text{ mol Al}} \times \frac{2g \text{ H}_2}{1 \text{ mol H}_2} = 12g \text{ H}_2$$

$$146g \text{ HCl} \times \frac{1 \text{ mol HCl}}{36.45g \text{ HCl}} \times \frac{3 \text{ mol H}_2}{6 \text{ mol HCl}} \times \frac{2g \text{ H}_2}{1 \text{ mol H}_2} = \boxed{4g \text{ H}_2}$$

HCl is limiting reactant

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