

Use your solubility curve chart to answer the following questions.

Part 1: Graph interpretation: (no work needed)

What are the customary units of solubility on solubility curves? **Grams of Solute per 100 g of water**

1. According to the graph, the solubility of any substance changes as **temperature** changes.
2. Most substances on this graph show increased solubility as temperature increases. What are the exceptions?  **$\text{Ca}_2(\text{SO}_4)_3$  and  $\text{NH}_3$**
3. The solubilities of substances whose curves show greater (steeper) slopes are **(more/less)** affected by temperature changes than those that have more gradual slopes.
4. Which salt has solubility values that are least affected by changes in temperature?  **$\text{NaCl}$**
5. What is the solubility of ammonium chloride at  $60^\circ\text{C}$ ? **50 g/100 g  $\text{H}_2\text{O}$**
6. At what temperature do potassium chlorate and potassium chloride have the same solubility in water? **95 °C**
7. Which compound is least soluble in water at  $12^\circ\text{C}$ ?  **$\text{KClO}_3$**
8. A saturated solution of which compound contains 130 grams of solute per 100 grams of water at  $70^\circ\text{C}$ ?  **$\text{KNO}_3$**
9. Are the following solutions unsaturated, saturated, or supersaturated?
  - a. 80 g of sodium nitrate in 100 g of water at  $30^\circ\text{C}$ . **unsaturated**
  - b. 80 g of potassium chlorate in 100 g of water at  $50^\circ\text{C}$ . **supersaturated**
10. Which saturated solution of a chloride has the greatest percentage by mass of solute at  $60^\circ\text{C}$ ? **don't worry about this question for the test**

Problems: Show all work on a separate sheet.

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**G. Marc Loudon, Joseph G. Stowell**

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