

- i. State the relative atomic mass of krypton
- ii. Calculate the density of krypton at r.t.p.

Answer:

- a) Helium has two electrons in its outermost shell while the other Group 0 elements have eight electrons in their outermost shells. All the Group 0 elements have stable electronic configurations. Thus these elements are uncreative and seldom react with other substances. In addition, these elements have very low melting and boiling point and exist as gases under room conditions.
- b) The boiling points of the elements increase down the group.
- c) Helium has a very low density and is unreactive (not flammable)
- d) i) relative mass of krypton is 84
 ii) Density of krypton at r.t.p. = $84 / 24 = 3.5 \text{ g/dm}^3$.

14.) The table below shows the physical properties of some transition metals.

Metal	Density gm^{-3}	Melting point / $^{\circ}\text{C}$	Boiling point / $^{\circ}\text{C}$	Relative electrical conductivity
Titanium	4.50	1660	3287	1
Vanadium	5.96	1890	3380	1.5
Chromium	7.20	1244	1962	5
Manganese	7.20	1244	1962	1
Iron	7.86	1535	2750	8.5
Cobalt	8.90	1495	2870	12
Nickel	8.90	1455	2730	12
Copper	8.92	1083	2567	49

- a) List two differences between the physical properties of Group I metals and the transition metals.

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