

# Unit 6 Study Guide

## TOPIC 1: REVIEW

Aq (aqueous) - dissolved in water

Solution: Homogenous mixture; solutes dissolved in solvents

Solute: Dissolved particles in a solution (i.e. NaCl)

Solvent: The dissolved medium in a solution



Saturated: Solution containing max amount of solute for a given amount of solvent at a constant temperature and pressure

## TOPIC 2: SOLUBILITY

- Water molecule has  $\ominus$  and  $\oplus$  end
- negative attached to cation
- positive attached to anion



Ion solution: water surrounding ions

- When certain IONIC COMPOUNDS (compounds w/ 1+ atoms) are added to water, they will break apart / DISSOCIATE



Solubility rules:

# Chemistry Unit 6 Study Guide

**Clemens Wendtner**



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## Table of Contents Chemistry Unit 6 Study Guide

1. Understanding the eBook Chemistry Unit 6 Study Guide
  - The Rise of Digital Reading Chemistry Unit 6 Study Guide
  - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Unit 6 Study Guide
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Chemistry Unit 6 Study Guide
  - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry Unit 6 Study Guide
  - Personalized Recommendations
  - Chemistry Unit 6 Study Guide User Reviews and Ratings
  - Chemistry Unit 6 Study Guide and Bestseller Lists
5. Accessing Chemistry Unit 6 Study Guide Free and Paid eBooks
  - Chemistry Unit 6 Study Guide Public Domain eBooks
  - Chemistry Unit 6 Study Guide eBook Subscription Services
  - Chemistry Unit 6 Study Guide Budget-Friendly Options
6. Navigating Chemistry Unit 6 Study Guide eBook Formats
  - ePub, PDF, MOBI, and More
  - Chemistry Unit 6 Study Guide Compatibility with Devices
  - Chemistry Unit 6 Study Guide Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Chemistry Unit 6 Study Guide
  - Highlighting and Note-Taking Chemistry Unit 6 Study Guide
  - Interactive Elements Chemistry Unit 6 Study Guide
8. Staying Engaged with Chemistry Unit 6 Study Guide

- Joining Online Reading Communities
- Participating in Virtual Book Clubs
- Following Authors and Publishers Chemistry Unit 6 Study Guide
- 9. Balancing eBooks and Physical Books Chemistry Unit 6 Study Guide
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Chemistry Unit 6 Study Guide
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Chemistry Unit 6 Study Guide
  - Setting Reading Goals Chemistry Unit 6 Study Guide
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chemistry Unit 6 Study Guide
  - Fact-Checking eBook Content of Chemistry Unit 6 Study Guide
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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web procedure 1 find the mass of a clean dry 100 ml beaker and record 2 measure out 8 00 grams of copper ii sulfate onto a piece of filter paper taking the filter paper s mass into account add these to the beaker 3 measure 50 0 ml of water in the

cylinder and add it to the beaker 4

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web description great lab for use during stoichiometry unit in chemistry reinforces ideas of mole ratios and stoichiometry calculations by measuring the mass of iron that reacts and the mass of copper metal produced students can calculate the ratio of moles of reactant to moles of product this experimental mole ratio from lab can be compared

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web in this experiment we will use stoichiometric principles to deduce the appropriate equation for the reaction between metallic iron and a solution of copper ii sulfate this reaction produces metallic copper which is seen precipitating as a finely divided red powder

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web through this experiment we sought to use stoichiometric principles which derive from stoichiometry the study of substance quantities in ratio or proportion in a chemical reaction in this lab experiment the principles help distinguish the equation between iron powder and aqueous copper ii sulfate in the iron copper ii sulfate redox reaction

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web this experiment is focused on a reaction between iron metal and a solution of copper ii sulfate we are able to figure out the chemical equation between the two by using stoichiometric principles the iron displaces the copper ii sulfate ions the copper ii is then reduced and gains electrons

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web stoichiometry can be defined as the relationship between the relative quantities of substances taking part in a reaction or forming a compound stoichiometric principles can be used to deduce the appropriate chemical equation for a reaction in our experiment we used solid iron in the form of powder was added to a solution of copper ii sulfate

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web under the finalize of the experiment it was found that iron was the limiting reactant and iron ii sulfate was a product of the reaction between 7 0002g away cooper ii sulfate and 2 0101g iron metal which produced 2 3037g of copper product with a percent yield of 100 72 introduction

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web stoichiometry lab report the reaction of iron with copper ii sulfate period 5 in this experiment i observed the reaction of iron with copper ii sulfate i obtained the results to this experiment through adding heated  $\text{CuSO}_4$  solution to the iron powder stirring the solution adding water then removing water adding acetone heating the acetone on low

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web stoichiometry the reaction of iron with copper ii sulfate purpose to enhance the understanding of stoichiometry a reaction between iron and copper ii sulfate solution will be conducted this will help you to differentiate limiting and excess reactant in a chemical reaction

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web stoichiometry lab iron and copper sulfate sample data data and observations mass of copper ii sulfate pentahydrate 3 27 g of g mass of filter paper mass of filter paper and dry copper mass of dry copper 1 28 g