

Post-lab questions:

1. Using the volumes of HSO_3^- solution added for each of the three titrations you conducted (A, B, and C), calculate the number of moles HSO_3^- required to reach the endpoint in each titration.
2. Using the moles of HSO_3^- calculated in #1 above, calculate the number of moles of MnO_4^- titrated according to the balanced equations you submitted in the prelab exercises.
3. Using the moles of MnO_4^- you calculated in #2 above, calculate the starting concentration of permanganate ion for each of the three titrations.
4. Write the electron configurations for Mn, Mn^{2+} , Mn^{4+} , Mn^{6+} , and Mn^{7+} .
5. Sodium bismuthate (NaBiO_3) oxidizes MnCl_2 to MnO_4^- and is consequently reduced to Bi(OH)_3 . How many grams of MnCl_2 will be oxidized by 7.5 g of NaBiO_3 ?
6. Balance the following oxidation-reduction reactions:
 - a. $\text{Mn}^{2+} + \text{BiO}_3^- \longrightarrow \text{MnO}_4^- + \text{Bi}^{3+}$ (in acid)
 - b. $\text{NO}_3^- + \text{I}_2 \longrightarrow \text{NO}_2 + \text{IO}_3^-$ (in acid)

- c. $\text{ClO}_3^- + \text{S}^{2-} \longrightarrow \text{Cl}^- + \text{S}$ (in base)
- d. $\text{Mo(OH)}_3 + \text{HO}_2^- \longrightarrow \text{MoO}_4^{2-}$ (in base)

- 7. In each of the titrations performed in this experiment, the end point was apparent due to a color change in the material being oxidized or reduced. Suppose that none of the materials involved in the redox reaction are colored. Suggest a way one could determine the endpoint in such a case. (HINT: Think about acid-base titrations.)

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