

Post-lab questions:

1. Using the volumes of HSO_3^- solution added for each of the three titrations you conducted (A, B, and C), calculate the number of moles HSO_3^- required to reach the endpoint in each titration.
 2. Using the moles of HSO_3^- calculated in #1 above, calculate the number of moles of MnO_4^- titrated according to the balanced equations you submitted in the prelab exercises.
 3. Using the moles of MnO_4^- you calculated in #2 above, calculate the starting concentration of permanganate ion for each of the three titrations.
 4. Write the electron configurations for Mn, Mn^{2+} , Mn^{4+} , Mn^{6+} , and Mn^{7+} .
 5. Sodium bismuthate (NaBiO_3) oxidizes MnCl_2 to MnO_4^- and is consequently reduced to Bi(OH)_3 . How many grams of MnCl_2 will be oxidized by 7.5 g of NaBiO_3 ?
 6. Balance the following oxidation-reduction reactions:
 - a. $\text{Mn}^{2+} + \text{BiO}_3^- \longrightarrow \text{MnO}_4^- + \text{Bi}^{3+}$ (in acid)
 - b. $\text{NO}_3^- + \text{I}_2 \longrightarrow \text{NO}_2 + \text{IO}_3^-$ (in acid)
-
- c. $\text{ClO}_3^- + \text{S}^{2-} \longrightarrow \text{Cl}^- + \text{S}$ (in base)
 - d. $\text{Mo(OH)}_3 + \text{HO}_2^- \longrightarrow \text{MoO}_4^{2-}$ (in base)
7. In each of the titrations performed in this experiment, the end point was apparent due to a color change in the material being oxidized or reduced. Suppose that none of the materials involved in the redox reaction are colored. Suggest a way one could determine the endpoint in such a case. (HINT: Think about acid-base titrations.)

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1. Understanding the eBook Conclusion Questions And Calculations Concentration And Molarity Post Lab Exercises Answers
 - The Rise of Digital Reading Conclusion Questions And Calculations Concentration And Molarity Post Lab Exercises Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Conclusion Questions And Calculations Concentration And Molarity Post Lab Exercises Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Conclusion Questions And Calculations Concentration And Molarity Post Lab Exercises Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Conclusion Questions And Calculations Concentration And Molarity Post Lab

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- Personalized Recommendations
 - Conclusion Questions And Calculations Concentration And Molarity Post Lab Exercises Answers User Reviews and Ratings
 - Conclusion Questions And Calculations Concentration And Molarity Post Lab Exercises Answers and Bestseller Lists
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 - Conclusion Questions And Calculations Concentration And Molarity Post Lab Exercises Answers Budget-Friendly Options
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 - Conclusion Questions And Calculations Concentration And Molarity Post Lab Exercises Answers Compatibility with Devices
 - Conclusion Questions And Calculations Concentration And Molarity Post Lab Exercises Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
- Adjustable Fonts and Text Sizes of Conclusion Questions And Calculations Concentration And Molarity Post Lab Exercises Answers
 - Highlighting and Note-Taking Conclusion Questions And Calculations Concentration And Molarity Post Lab Exercises Answers
 - Interactive Elements Conclusion Questions And Calculations Concentration And Molarity Post Lab Exercises Answers
8. Staying Engaged with Conclusion Questions And Calculations Concentration And Molarity Post Lab Exercises Answers
- Joining Online Reading Communities

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- Participating in Virtual Book Clubs
- Following Authors and Publishers Conclusion Questions And Calculations Concentration And Molarity Post Lab Exercises Answers
- 9. Balancing eBooks and Physical Books Conclusion Questions And Calculations Concentration And Molarity Post Lab Exercises Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Conclusion Questions And Calculations Concentration And Molarity Post Lab Exercises Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Conclusion Questions And Calculations Concentration And Molarity Post Lab Exercises Answers
 - Setting Reading Goals Conclusion Questions And Calculations Concentration And Molarity Post Lab Exercises Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Conclusion Questions And Calculations Concentration And Molarity Post Lab Exercises Answers
 - Fact-Checking eBook Content of Conclusion Questions And Calculations Concentration And Molarity Post Lab Exercises Answers
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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