

Molecular Models Lab Covalent Compounds

Introduction

The purpose of this lab is to evaluate the geometries and polarities of simple covalent molecules. Many physical and chemical properties of compounds depend upon the shape or geometry of the molecule and its polarity. A hemoglobin molecule, for example, becomes ineffective as an oxygen transport vehicle if its shape is altered. Molecular shape determines a compound's boiling point, freezing point, and viscosity. A molecule's polarity determines whether it dissolves in water or not.

Molecular Geometries

The geometry of a small molecule can be predicted by examining the central atom. The valence electron pairs that surround the central atom will repel each other. The electron pairs will position themselves so that there is a maximum distance between them. This rule is called the **VSEPR** theory (Valence Shell Electron Pair Repulsion). The electron pairs include the covalent bonds (shared pairs) and the unshared pairs of electrons. The 5 simplest geometries and their characteristics are found in the following table:

| No. of Electron Pairs on Central Atom | Geometry Name | Central Atom Characteristics | Comments |
|---------------------------------------|-----------------|--|---|
| No central atom | Linear | $X - X$ | If a molecule only contains 2 atoms it is linear. |
| 2 shared, 0 unshared | Linear | $— X —$ or $= X =$ | Double and triple bonds only count as a shared pair for geometry purposes. |
| 3 shared, 0 unshared | Trigonal planar | $\begin{array}{c} \\ — X — \end{array}$ or $\begin{array}{c} \\ = X = \end{array}$ | Bond angles are 120° |
| 4 shared, 0 unshared | Tetrahedral | $\begin{array}{c} \\ — X — \\ \end{array}$ | Since the molecule is 3-dimensional, the bond angles are 109° , not 90° . |
| 2 shared, 2 unshared | Bent | $\begin{array}{c} \\ : X : \\ \end{array}$ | The bond angle is slightly $< 109^\circ$, not 90° . |
| 3 shared, 1 unshared | Pyramidal | $\begin{array}{c} \\ — X : \\ \end{array}$ | The bond angle is slightly $< 109^\circ$, not 90° . |

Conclusion For Molecular Models Lab In Chemistry

**Dakeshwar Kumar
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Aslam**



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Exploring General Chemistry in the Laboratory Colleen F. Craig, Kim N. Gunnerson, 2017-02-01 This laboratory manual is intended for a two semester general chemistry course The procedures are written with the goal of simplifying a complicated and often challenging subject for students by applying concepts to everyday life This lab manual covers topics such as composition of compounds reactivity stoichiometry limiting reactants gas laws calorimetry periodic trends molecular structure spectroscopy kinetics equilibria thermodynamics electrochemistry intermolecular forces solutions and coordination complexes By the end of this course you should have a solid understanding of the basic concepts of chemistry which will give you confidence as you embark on your career in science

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