

## Chapter 2 The Nature of Matter • Section 1 Summary

## Describing Matter

### Key Concepts

- What kinds of properties are used to describe matter?
- What are elements, and how do they relate to compounds?
- What are the properties of a mixture?

**Matter** is anything that has mass and takes up space. **Chemistry** is the study of the properties of matter and how matter changes. In chemistry, a **substance** is a single kind of matter that is pure.

Every form of matter has two kinds of properties—**physical properties** and **chemical properties**. A **physical property** is observed without changing a substance into another substance. Examples of physical properties are hardness, texture, color, and ability to dissolve in water. A **chemical property** is the ability of a substance to change into different substances. Some chemical properties are burning and rusting.

All matter is made up of elements. An **element** is a pure substance that cannot be broken down into any other substance. **Elements** are the simplest substances. Each element is identified by its specific physical and chemical properties. An **atom** is the basic particle that makes up an element. Atoms of most elements can combine with other atoms. A **chemical bond** is the force that holds two atoms together. Atoms often combine to form **molecules**, which are larger particles made of two or more atoms held together by chemical bonds.

When elements are chemically combined, they form **compounds** having properties that are different from those of the uncombined elements. A **compound** is a pure substance made of two or more elements chemically combined in a set ratio. A compound may be represented by a **chemical formula**. A chemical formula shows the elements in the compound and the ratio of atoms. For example, the chemical formula for carbon dioxide is  $\text{CO}_2$ . In carbon dioxide, there are always two oxygen atoms to every one carbon atom.

Elements and compounds are pure substances, but most of the materials you see every day are not. Instead, they are mixtures. A **mixture** is made of two or more substances that are together in the same place, but are not chemically combined. Mixtures differ from compounds in two ways. Each substance in a mixture keeps its individual properties. Also, the parts of a mixture are not combined in a set ratio.

A mixture can be heterogeneous or homogeneous. In a **heterogeneous mixture**, you can see the different parts. The substances in a **homogeneous mixture** are so evenly mixed that you cannot see the different parts. A **solution** is an example of a homogeneous mixture. Air is a solution of nitrogen gas, oxygen gas, plus small amounts of other gases. Unlike compounds, mixtures are easily separated into their components. For example, iron filings can be easily removed from salt with a magnet.

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